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## Structure Reports

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# Diaquabis(2-chlorobenzoato- $\kappa$ O)bis-(*N,N*-diethylnicotinamide- $\kappa$ N<sup>1</sup>)nickel(II)

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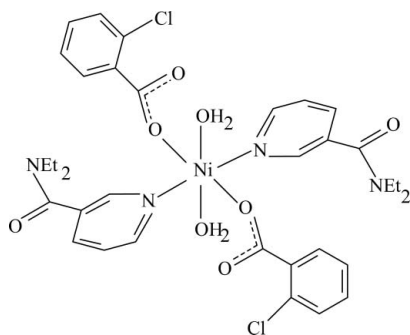
 Key indicators: single-crystal X-ray study;  $T = 100$  K; mean  $\sigma(\text{C}-\text{C}) = 0.002$  Å;

 $R$  factor = 0.031;  $wR$  factor = 0.084; data-to-parameter ratio = 19.4.

In the monomeric and centrosymmetric title Ni<sup>II</sup> complex,  $[\text{Ni}(\text{C}_7\text{H}_4\text{ClO}_2)_2(\text{C}_{10}\text{H}_{14}\text{N}_2\text{O})_2(\text{H}_2\text{O})_2]$ , the Ni<sup>II</sup> ion is located on an inversion center. The asymmetric unit contains one 2-chlorobenzoate ligand, one diethylnicotinamide (DENA) ligand and one coordinating water molecule, the ligands being monodentate. The four O atoms in the equatorial plane around the Ni<sup>II</sup> ion form a slightly distorted square-planar arrangement, while the slightly distorted octahedral coordination is completed by two N atoms of the DENA ligands in axial positions. The dihedral angle between the benzene ring and the attached carboxylate group is  $87.36(10)^\circ$ , while the pyridine and benzene rings are oriented at an angle of  $41.90(5)^\circ$ . In the crystal structure, intermolecular O—H...O hydrogen bonds link the molecules into a two-dimensional network parallel to  $(10\bar{1})$ .

## Related literature

For general background, see: Antolini *et al.* (1982); Bigoli *et al.* (1972); Nadzhafov *et al.* (1981); Shnulin *et al.* (1981). For related structures, see: Hökelek *et al.* (1995, 1997, 2007, 2008); Hökelek & Necefoğlu (1996, 1997, 2007).



## Experimental

### Crystal data

 $[\text{Ni}(\text{C}_7\text{H}_4\text{ClO}_2)_2(\text{C}_{10}\text{H}_{14}\text{N}_2\text{O})_2(\text{H}_2\text{O})_2]$ 
 $M_r = 762.31$ 

 Monoclinic,  $P2_1/n$ 
 $a = 12.7505(2)$  Å

 $b = 10.3565(2)$  Å

 $c = 14.9673(3)$  Å

 $\beta = 114.046(1)^\circ$ 
 $V = 1804.92(6)$  Å<sup>3</sup>
 $Z = 2$ 

 Mo  $K\alpha$  radiation

 $\mu = 0.74$  mm<sup>-1</sup>
 $T = 100$  K

 $0.27 \times 0.18 \times 0.11$  mm

### Data collection

Bruker Kappa APEXII CCD area-detector diffractometer

Absorption correction: multi-scan (SADABS; Bruker, 2005)

 $T_{\text{min}} = 0.828$ ,  $T_{\text{max}} = 0.923$ 

16593 measured reflections

4519 independent reflections

 3781 reflections with  $I > 2\sigma(I)$ 
 $R_{\text{int}} = 0.034$ 

### Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.031$ 
 $wR(F^2) = 0.084$ 
 $S = 1.06$ 

4519 reflections

233 parameters

2 restraints

H atoms treated by a mixture of independent and constrained refinement

 $\Delta\rho_{\text{max}} = 1.08$  e Å<sup>-3</sup>
 $\Delta\rho_{\text{min}} = -0.39$  e Å<sup>-3</sup>
**Table 1**

Selected bond lengths (Å).

Ni1—O1	2.0336 (10)	Ni1—N1	2.1181 (12)
Ni1—O4	2.0867 (10)		

**Table 2**

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O4—H41...O2 <sup>i</sup>	0.82 (2)	1.86 (2)	2.6267 (17)	155 (2)
O4—H42...O3 <sup>ii</sup>	0.85 (2)	1.93 (2)	2.7826 (15)	172 (2)

 Symmetry codes: (i)  $-x, -y + 1, -z$ ; (ii)  $-x + \frac{1}{2}, y - \frac{1}{2}, -z + \frac{1}{2}$ .

Data collection: APEX2 (Bruker, 2007); cell refinement: SAINT (Bruker, 2007); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: ORTEP-3 for Windows (Farrugia, 1997); software used to prepare material for publication: WinGX (Farrugia, 1999).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: CI2776).

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## supporting information

*Acta Cryst.* (2009). E65, m545–m546 [doi:10.1107/S1600536809014226]

**Diaquabis(2-chlorobenzoato- $\kappa$ O)bis(*N,N*-diethylnicotinamide- $\kappa$ N<sup>1</sup>)nickel(II)****Tuncer Hökelek, Hakan Dal, Barış Tercan, F. Elif Özbek and Hacali Necefoğlu****S1. Comment**

Transition metal complexes with biochemically active ligands frequently show interesting physical and/or chemical properties, as a result they may find applications in biological systems (Antolini *et al.*, 1982). The structural functions and coordination relationships of the arylcarboxylate ion in transition metal complexes of benzoic acid derivatives change depending on the nature and position of the substituent groups on the benzene ring, the nature of the additional ligand molecule or solvent, and the medium of the synthesis (Nadzhafov *et al.*, 1981; Shnulin *et al.*, 1981). The nicotinic acid derivative *N,N*-diethylnicotinamide (DENA) is an important respiratory stimulant (Bigoli *et al.*, 1972).

The structure determination of the title compound, (I), a nickel complex with two 2-chlorobenzoate (CB), two diethylnicotinamide (DENA) ligands and two water molecules, was undertaken in order to determine the properties of the ligands and also to compare the results obtained with those reported previously.

Compound (I) is a monomeric complex, with the Ni<sup>II</sup> ion on a centre of symmetry. It contains two CB, two DENA ligands and two water molecules (Fig. 1). All ligands are monodentate. The four O atoms (O1, O4, and the symmetry-related atoms O1', O4') in the equatorial plane around the Ni<sup>II</sup> ion form a slightly distorted square-planar arrangement, while the slightly distorted octahedral coordination is completed by the two N atoms of the DENA ligands (N1, N1') in the axial positions (Table 1 and Fig. 1).

The near equality of the C1—O1 [1.2616 (17) Å] and C1—O2 [1.2435 (18) Å] bonds in the carboxylate group indicates a delocalized bonding arrangement, rather than localized single and double bonds, and may be compared with the corresponding distances: 1.256 (6) and 1.245 (6) Å in [Mn(DENA)<sub>2</sub>(C<sub>7</sub>H<sub>4</sub>ClO<sub>2</sub>)<sub>2</sub>(H<sub>2</sub>O)<sub>2</sub>], (II) (Hökelek *et al.*, 2008), 1.265 (6) and 1.275 (6) Å in [Mn(C<sub>9</sub>H<sub>10</sub>NO<sub>2</sub>)<sub>2</sub>(H<sub>2</sub>O)<sub>4</sub>].2(H<sub>2</sub>O), (III) (Hökelek & Necefoğlu, 2007), 1.260 (4) and 1.252 (4) Å in [Zn(DENA)<sub>2</sub>(C<sub>7</sub>H<sub>4</sub>FO<sub>2</sub>)<sub>2</sub>(H<sub>2</sub>O)<sub>2</sub>], (IV) (Hökelek *et al.*, 2007), 1.259 (9) and 1.273 (9) Å in Cu<sub>2</sub>(DENA)<sub>2</sub>(C<sub>6</sub>H<sub>5</sub>COO)<sub>4</sub>, (V) (Hökelek *et al.*, 1995), 1.279 (4) and 1.246 (4) Å in [Zn<sub>2</sub>(DENA)<sub>2</sub>(C<sub>7</sub>H<sub>5</sub>O<sub>3</sub>)<sub>4</sub>].2H<sub>2</sub>O, (VI) (Hökelek & Necefoğlu, 1996), 1.251 (6) and 1.254 (7) Å in [Co(DENA)<sub>2</sub>(C<sub>7</sub>H<sub>5</sub>O<sub>3</sub>)<sub>2</sub>(H<sub>2</sub>O)<sub>2</sub>], (VII) (Hökelek & Necefoğlu, 1997) and 1.278 (3) and 1.246 (3) Å in [Cu(DENA)<sub>2</sub>(C<sub>7</sub>H<sub>4</sub>NO<sub>4</sub>)<sub>2</sub>(H<sub>2</sub>O)<sub>2</sub>], (VIII) (Hökelek *et al.*, 1997). In (I), the average Ni—O bond length is 2.0602 (10) Å and the Ni1 atom is displaced out of the least-squares plane of the carboxylate group (O1/C1/O2) by -0.276 (1) Å. The dihedral angle between the planar carboxylate group and the benzene ring A (C2—C7) is 87.36 (10)°, while that between rings A and B (N1/C8—C12) is 41.90 (5)°.

In the crystal structure, intermolecular O—H...O hydrogen bonds (Table 2) link the molecules into a two-dimensional network parallel to the (1 0  $\bar{1}$ ).

**S2. Experimental**

The title compound was prepared by the reaction of Ni(SO<sub>4</sub>).6(H<sub>2</sub>O) (1.31 g, 5 mmol) in H<sub>2</sub>O (20 ml) and DENA (1.78 g, 10 mmol) in H<sub>2</sub>O (20 ml) with sodium 2-chlorobenzoate (1.785 g, 10 mmol) in H<sub>2</sub>O (50 ml). The mixture was filtered and set aside to crystallize at ambient temperature for 3 d, giving blue single crystals.

## S3. Refinement

H atoms of water molecule were located in a difference Fourier map and refined isotropically, with a O-H restraint. The remaining H atoms were positioned geometrically with C-H = 0.93, 0.97 and 0.96 Å, for aromatic, methylene and methyl H atoms and constrained to ride on their parent atoms, with  $U_{\text{iso}}(\text{H}) = xU_{\text{eq}}(\text{C})$ , where  $x = 1.5$  for methyl H and  $x = 1.2$  for all other H atoms.

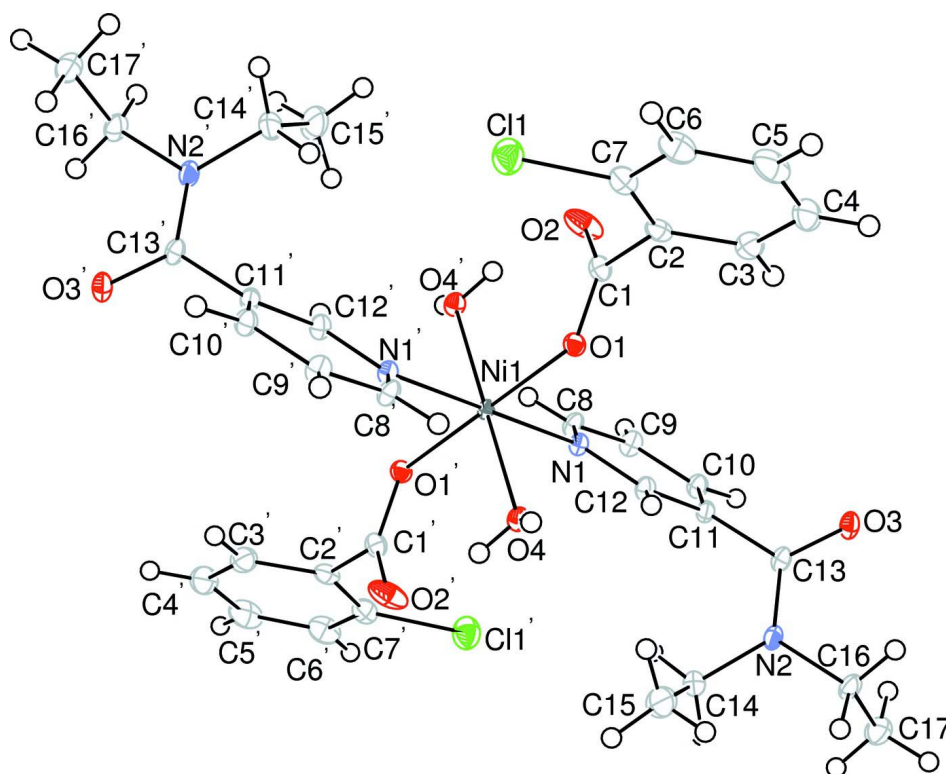


Figure 1

The molecular structure of the title compound with the atom-numbering scheme. Displacement ellipsoids are drawn at the 50% probability level. Primed atoms are generated by the symmetry operator  $(-x, 1-y, -z)$ .

**Diaquabis(2-chlorobenzoato- $\kappa$ O)bis(*N,N*-diethylnicotinamide- $\kappa$ N<sup>1</sup>)nickel(II)***Crystal data*

$[\text{Ni}(\text{C}_7\text{H}_4\text{ClO}_2)_2(\text{C}_{10}\text{H}_{14}\text{N}_2\text{O})_2(\text{H}_2\text{O})_2]$

$M_r = 762.31$

Monoclinic,  $P2_1/n$

Hall symbol:  $-P\ 2_1/n$

$a = 12.7505(2)\ \text{\AA}$

$b = 10.3565(2)\ \text{\AA}$

$c = 14.9673(3)\ \text{\AA}$

$\beta = 114.046(1)^\circ$

$V = 1804.92(6)\ \text{\AA}^3$

$Z = 2$

$F(000) = 796$

$D_x = 1.403\ \text{Mg m}^{-3}$

Mo  $K\alpha$  radiation,  $\lambda = 0.71073\ \text{\AA}$

Cell parameters from 7783 reflections

$\theta = 2.5\text{--}28.4^\circ$

$\mu = 0.74\ \text{mm}^{-1}$

$T = 100\ \text{K}$

Block, blue

$0.27 \times 0.18 \times 0.11\ \text{mm}$

*Data collection*

Bruker Kappa APEXII CCD area-detector  
diffractometer  
Radiation source: fine-focus sealed tube  
Graphite monochromator  
 $\varphi$  and  $\omega$  scans  
Absorption correction: multi-scan  
(*SADABS*; Bruker, 2005)  
 $T_{\min} = 0.828$ ,  $T_{\max} = 0.923$

16593 measured reflections  
4519 independent reflections  
3781 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.034$   
 $\theta_{\max} = 28.4^\circ$ ,  $\theta_{\min} = 1.8^\circ$   
 $h = -17 \rightarrow 17$   
 $k = -13 \rightarrow 12$   
 $l = -20 \rightarrow 18$

*Refinement*

Refinement on  $F^2$   
Least-squares matrix: full  
 $R[F^2 > 2\sigma(F^2)] = 0.031$   
 $wR(F^2) = 0.084$   
 $S = 1.06$   
4519 reflections  
233 parameters  
2 restraints  
Primary atom site location: structure-invariant  
direct methods

Secondary atom site location: difference Fourier  
map  
Hydrogen site location: inferred from  
neighbouring sites  
H atoms treated by a mixture of independent  
and constrained refinement  
 $w = 1/[\sigma^2(F_o^2) + (0.0456P)^2 + 0.2566P]$   
where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\max} = 0.001$   
 $\Delta\rho_{\max} = 1.08 \text{ e } \text{\AA}^{-3}$   
 $\Delta\rho_{\min} = -0.39 \text{ e } \text{\AA}^{-3}$

*Special details*

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Ni1	0.0000	0.5000	0.0000	0.00853 (8)
Cl1	-0.31050 (4)	0.35043 (4)	0.08471 (3)	0.02678 (11)
O1	-0.03038 (8)	0.44605 (10)	0.11794 (7)	0.0128 (2)
O2	-0.13227 (11)	0.61462 (11)	0.13322 (9)	0.0268 (3)
O3	0.43014 (9)	0.61598 (10)	0.39621 (7)	0.0151 (2)
O4	0.11786 (9)	0.34845 (10)	0.03504 (8)	0.0123 (2)
H41	0.1299 (19)	0.338 (2)	-0.0143 (13)	0.042 (6)*
H42	0.1100 (18)	0.2746 (17)	0.0570 (15)	0.043 (6)*
N1	0.13740 (10)	0.61957 (11)	0.08958 (9)	0.0114 (2)
N2	0.48379 (11)	0.49545 (12)	0.29580 (9)	0.0152 (3)
C1	-0.09075 (13)	0.50440 (14)	0.15415 (11)	0.0135 (3)
C2	-0.11211 (13)	0.42820 (14)	0.23115 (11)	0.0148 (3)
C3	-0.03393 (14)	0.42920 (16)	0.32862 (12)	0.0202 (3)
H3	0.0299	0.4830	0.3480	0.024*
C4	-0.05013 (15)	0.35049 (17)	0.39746 (12)	0.0259 (4)

H4	0.0027	0.3517	0.4624	0.031*
C5	-0.14508 (16)	0.27062 (17)	0.36887 (13)	0.0270 (4)
H5	-0.1551	0.2169	0.4145	0.032*
C6	-0.22513 (15)	0.27003 (16)	0.27308 (13)	0.0235 (4)
H6	-0.2895	0.2170	0.2541	0.028*
C7	-0.20830 (14)	0.34949 (15)	0.20560 (11)	0.0183 (3)
C8	0.14003 (12)	0.74683 (14)	0.07426 (11)	0.0136 (3)
H8	0.0804	0.7833	0.0209	0.016*
C9	0.22780 (12)	0.82654 (14)	0.13450 (11)	0.0151 (3)
H9	0.2255	0.9148	0.1223	0.018*
C10	0.31847 (12)	0.77378 (14)	0.21268 (11)	0.0139 (3)
H10	0.3776	0.8255	0.2547	0.017*
C11	0.31889 (12)	0.64090 (14)	0.22689 (10)	0.0119 (3)
C12	0.22679 (12)	0.56819 (14)	0.16507 (10)	0.0124 (3)
H12	0.2267	0.4799	0.1762	0.015*
C13	0.41594 (12)	0.58190 (13)	0.31243 (11)	0.0120 (3)
C14	0.47757 (15)	0.45998 (18)	0.19868 (12)	0.0254 (4)
H14A	0.5535	0.4670	0.1989	0.031*
H14B	0.4278	0.5205	0.1507	0.031*
C15	0.43258 (16)	0.32361 (19)	0.16849 (14)	0.0368 (5)
H15A	0.4372	0.3023	0.1078	0.055*
H15B	0.3541	0.3187	0.1605	0.055*
H15C	0.4781	0.2638	0.2181	0.055*
C16	0.58189 (13)	0.44280 (15)	0.38015 (11)	0.0166 (3)
H16A	0.5967	0.3551	0.3657	0.020*
H16B	0.5628	0.4401	0.4366	0.020*
C17	0.68933 (14)	0.52372 (17)	0.40425 (13)	0.0233 (4)
H17A	0.7512	0.4876	0.4600	0.035*
H17B	0.6750	0.6105	0.4188	0.035*
H17C	0.7098	0.5242	0.3492	0.035*

Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Ni1	0.00850 (13)	0.00775 (13)	0.00768 (13)	0.00027 (9)	0.00159 (10)	0.00033 (9)
Cl1	0.0248 (2)	0.0307 (2)	0.0230 (2)	-0.00606 (17)	0.00789 (18)	-0.00259 (17)
O1	0.0143 (5)	0.0135 (5)	0.0110 (5)	0.0022 (4)	0.0055 (4)	0.0018 (4)
O2	0.0457 (8)	0.0154 (6)	0.0323 (7)	0.0119 (5)	0.0293 (6)	0.0083 (5)
O3	0.0180 (5)	0.0122 (5)	0.0103 (5)	0.0014 (4)	0.0008 (4)	-0.0013 (4)
O4	0.0136 (5)	0.0093 (5)	0.0121 (5)	0.0011 (4)	0.0034 (4)	0.0006 (4)
N1	0.0114 (6)	0.0109 (6)	0.0102 (6)	-0.0004 (5)	0.0027 (5)	0.0001 (5)
N2	0.0143 (6)	0.0175 (7)	0.0098 (6)	0.0038 (5)	0.0009 (5)	0.0004 (5)
C1	0.0158 (7)	0.0125 (7)	0.0126 (7)	-0.0021 (6)	0.0061 (6)	-0.0011 (6)
C2	0.0181 (7)	0.0132 (7)	0.0166 (7)	0.0037 (6)	0.0106 (6)	0.0016 (6)
C3	0.0185 (8)	0.0232 (9)	0.0194 (8)	0.0028 (6)	0.0081 (7)	0.0026 (7)
C4	0.0288 (9)	0.0322 (10)	0.0175 (8)	0.0105 (8)	0.0104 (7)	0.0082 (7)
C5	0.0380 (10)	0.0253 (9)	0.0269 (9)	0.0072 (8)	0.0228 (9)	0.0114 (7)
C6	0.0293 (9)	0.0199 (9)	0.0286 (9)	-0.0026 (7)	0.0193 (8)	0.0018 (7)

C7	0.0212 (8)	0.0181 (8)	0.0175 (8)	0.0008 (6)	0.0098 (7)	0.0002 (6)
C8	0.0125 (7)	0.0120 (7)	0.0131 (7)	0.0019 (6)	0.0018 (6)	0.0028 (6)
C9	0.0169 (7)	0.0093 (7)	0.0164 (7)	-0.0002 (6)	0.0039 (6)	0.0010 (6)
C10	0.0137 (7)	0.0125 (7)	0.0133 (7)	-0.0022 (6)	0.0033 (6)	-0.0019 (6)
C11	0.0117 (7)	0.0120 (7)	0.0097 (7)	0.0005 (5)	0.0022 (6)	-0.0006 (5)
C12	0.0144 (7)	0.0097 (7)	0.0113 (7)	0.0006 (5)	0.0032 (6)	0.0012 (5)
C13	0.0114 (7)	0.0089 (7)	0.0118 (7)	-0.0023 (5)	0.0009 (6)	-0.0005 (5)
C14	0.0223 (9)	0.0383 (10)	0.0129 (8)	0.0119 (7)	0.0042 (7)	-0.0016 (7)
C15	0.0259 (10)	0.0442 (12)	0.0291 (10)	0.0087 (8)	-0.0003 (8)	-0.0213 (9)
C16	0.0159 (7)	0.0144 (8)	0.0147 (7)	0.0054 (6)	0.0012 (6)	0.0021 (6)
C17	0.0169 (8)	0.0271 (9)	0.0214 (9)	0.0006 (7)	0.0033 (7)	-0.0046 (7)

*Geometric parameters (Å, °)*

Ni1—O1 <sup>i</sup>	2.0336 (10)	C7—C2	1.390 (2)
Ni1—O1	2.0336 (10)	C7—C6	1.386 (2)
Ni1—O4	2.0867 (10)	C8—C9	1.387 (2)
Ni1—O4 <sup>i</sup>	2.0867 (10)	C8—H8	0.93
Ni1—N1 <sup>i</sup>	2.1181 (12)	C9—H9	0.93
Ni1—N1	2.1181 (12)	C10—C9	1.379 (2)
C11—C7	1.7466 (17)	C10—H10	0.93
O1—C1	1.2616 (17)	C11—C10	1.392 (2)
O2—C1	1.2435 (18)	C11—C12	1.384 (2)
O3—C13	1.2429 (17)	C12—H12	0.93
O4—H41	0.820 (15)	C13—N2	1.3369 (18)
O4—H42	0.854 (15)	C13—C11	1.501 (2)
N1—C8	1.3405 (19)	C14—H14A	0.97
N1—C12	1.3452 (18)	C14—H14B	0.97
N2—C14	1.470 (2)	C15—C14	1.522 (3)
N2—C16	1.4728 (19)	C15—H15A	0.96
C2—C3	1.392 (2)	C15—H15B	0.96
C2—C1	1.510 (2)	C15—H15C	0.96
C3—C4	1.394 (2)	C16—H16A	0.97
C3—H3	0.93	C16—H16B	0.97
C4—H4	0.93	C17—C16	1.519 (2)
C5—C4	1.382 (3)	C17—H17A	0.96
C5—H5	0.93	C17—H17B	0.96
C6—C5	1.380 (2)	C17—H17C	0.96
C6—H6	0.93		
O1 <sup>i</sup> —Ni1—O1	180.0	C6—C7—C2	121.94 (15)
O1 <sup>i</sup> —Ni1—O4	93.00 (4)	C6—C7—C11	119.13 (13)
O1—Ni1—O4	87.00 (4)	N1—C8—C9	122.93 (14)
O1 <sup>i</sup> —Ni1—O4 <sup>i</sup>	87.00 (4)	N1—C8—H8	118.5
O1—Ni1—O4 <sup>i</sup>	93.00 (4)	C9—C8—H8	118.5
O1 <sup>i</sup> —Ni1—N1 <sup>i</sup>	90.70 (4)	C8—C9—H9	120.3
O1—Ni1—N1 <sup>i</sup>	89.30 (4)	C10—C9—C8	119.41 (14)
O1 <sup>i</sup> —Ni1—N1	89.30 (4)	C10—C9—H9	120.3

O1—Ni1—N1	90.70 (4)	C9—C10—C11	118.10 (13)
O4—Ni1—O4 <sup>i</sup>	180.0	C9—C10—H10	121.0
O4—Ni1—N1 <sup>i</sup>	92.55 (4)	C11—C10—H10	121.0
O4 <sup>i</sup> —Ni1—N1 <sup>i</sup>	87.45 (4)	C10—C11—C13	118.90 (13)
O4—Ni1—N1	87.45 (4)	C12—C11—C10	119.09 (13)
O4 <sup>i</sup> —Ni1—N1	92.55 (4)	C12—C11—C13	121.91 (13)
N1 <sup>i</sup> —Ni1—N1	180.00 (7)	N1—C12—C11	122.94 (13)
C1—O1—Ni1	127.35 (9)	N1—C12—H12	118.5
Ni1—O4—H41	104.4 (15)	C11—C12—H12	118.5
Ni1—O4—H42	125.8 (15)	O3—C13—N2	122.60 (13)
H41—O4—H42	109 (2)	O3—C13—C11	118.39 (13)
C8—N1—Ni1	122.58 (10)	N2—C13—C11	119.01 (13)
C8—N1—C12	117.45 (12)	N2—C14—C15	112.80 (15)
C12—N1—Ni1	119.97 (9)	N2—C14—H14A	109.0
C13—N2—C14	124.98 (13)	N2—C14—H14B	109.0
C13—N2—C16	118.44 (12)	C15—C14—H14A	109.0
C14—N2—C16	116.10 (12)	C15—C14—H14B	109.0
O1—C1—C2	114.14 (12)	H14A—C14—H14B	107.8
O2—C1—O1	127.09 (14)	C14—C15—H15A	109.5
O2—C1—C2	118.77 (13)	C14—C15—H15B	109.5
C3—C2—C1	121.41 (14)	C14—C15—H15C	109.5
C7—C2—C1	120.65 (14)	H15A—C15—H15B	109.5
C7—C2—C3	117.87 (14)	H15A—C15—H15C	109.5
C2—C3—C4	120.80 (16)	H15B—C15—H15C	109.5
C2—C3—H3	119.6	N2—C16—C17	111.60 (13)
C4—C3—H3	119.6	N2—C16—H16A	109.3
C3—C4—H4	120.1	N2—C16—H16B	109.3
C5—C4—C3	119.74 (16)	C17—C16—H16A	109.3
C5—C4—H4	120.1	C17—C16—H16B	109.3
C6—C5—C4	120.52 (15)	H16A—C16—H16B	108.0
C6—C5—H5	119.7	C16—C17—H17A	109.5
C4—C5—H5	119.7	C16—C17—H17B	109.5
C5—C6—C7	119.08 (16)	C16—C17—H17C	109.5
C5—C6—H6	120.5	H17A—C17—H17B	109.5
C7—C6—H6	120.5	H17A—C17—H17C	109.5
C2—C7—C11	118.94 (12)	H17B—C17—H17C	109.5
O4—Ni1—O1—C1	172.55 (12)	C1—C2—C3—C4	-174.81 (14)
O4 <sup>i</sup> —Ni1—O1—C1	-7.45 (12)	C7—C2—C3—C4	2.0 (2)
N1 <sup>i</sup> —Ni1—O1—C1	-94.85 (12)	C2—C3—C4—C5	-0.2 (2)
N1—Ni1—O1—C1	85.15 (12)	C6—C5—C4—C3	-1.3 (3)
O1 <sup>i</sup> —Ni1—N1—C8	58.61 (11)	C7—C6—C5—C4	0.8 (3)
O1—Ni1—N1—C8	-121.39 (11)	C11—C7—C2—C1	-5.36 (19)
O1 <sup>i</sup> —Ni1—N1—C12	-121.15 (11)	C11—C7—C2—C3	177.78 (12)
O1—Ni1—N1—C12	58.85 (11)	C6—C7—C2—C1	174.35 (14)
O4—Ni1—N1—C8	151.65 (11)	C6—C7—C2—C3	-2.5 (2)
O4 <sup>i</sup> —Ni1—N1—C8	-28.35 (11)	C11—C7—C6—C5	-179.16 (12)
O4—Ni1—N1—C12	-28.11 (11)	C2—C7—C6—C5	1.1 (2)



O4 <sup>i</sup> —Ni1—N1—C12	151.89 (11)	N1—C8—C9—C10	1.6 (2)
Ni1—O1—C1—O2	-9.8 (2)	C11—C10—C9—C8	1.1 (2)
Ni1—O1—C1—C2	170.38 (9)	C12—C11—C10—C9	-2.7 (2)
Ni1—N1—C8—C9	177.78 (11)	C13—C11—C10—C9	-179.11 (13)
C12—N1—C8—C9	-2.5 (2)	C10—C11—C12—N1	1.8 (2)
Ni1—N1—C12—C11	-179.49 (11)	C13—C11—C12—N1	178.16 (13)
C8—N1—C12—C11	0.7 (2)	O3—C13—N2—C14	-174.33 (15)
C13—N2—C14—C15	-110.16 (17)	O3—C13—N2—C16	-2.5 (2)
C16—N2—C14—C15	77.88 (18)	C11—C13—N2—C14	5.4 (2)
C13—N2—C16—C17	-90.03 (17)	C11—C13—N2—C16	177.22 (12)
C14—N2—C16—C17	82.48 (17)	O3—C13—C11—C10	61.59 (19)
C3—C2—C1—O1	85.51 (18)	O3—C13—C11—C12	-114.75 (16)
C3—C2—C1—O2	-94.30 (19)	N2—C13—C11—C10	-118.19 (15)
C7—C2—C1—O1	-91.23 (17)	N2—C13—C11—C12	65.48 (19)
C7—C2—C1—O2	88.96 (19)		

Symmetry code: (i)  $-x, -y+1, -z$ .

*Hydrogen-bond geometry (Å, °)*

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
O4—H41 $\cdots$ O2 <sup>i</sup>	0.82 (2)	1.86 (2)	2.6267 (17)	155 (2)
O4—H42 $\cdots$ O3 <sup>ii</sup>	0.85 (2)	1.93 (2)	2.7826 (15)	172 (2)

Symmetry codes: (i)  $-x, -y+1, -z$ ; (ii)  $-x+1/2, y-1/2, -z+1/2$ .